

Chapter 14

Give the four factors that affect a reaction rate. For each give a brief explanation why the factor affects the rate.

Explain a reaction and reaction rate on a molecular level-

What is the usual unit for a rate of reaction?

What are two ways to express a reaction rate?

Why is a negative sign used when using the rate of disappearance?

Work through sample exercise 14.1-

How is an instantaneous rate determined for a reaction?

In your own words explain how Stoichiometry applies to reaction rates.

Give the general rate law equation-

What is k ?

What do the exponents in a rate law equation tell us?

If a rate law is second order with respect to a certain reactant how will doubling the concentration of the reactant affect the rate?

How are units for the rate constant derived?

How is the rate law determined?

Work through sample exercise 14.6-