

The concepts in chapter 4 are pivotal to your success in this class and on the AP test. Make sure you keep track of homework due dates and keep on top of the material. It is not the most difficult material, but it is the most important so far in this class, if not THE most important material in the entire year.

Read sections 4.1 through 4.2 to answer the questions below.

How is the solute and solvent identified in a solution?

Explain why a solution of salt will conduct and a solution of sucrose will not-

Based on the reading what is one way to determine between an Ionic and a non-ionic solution?

In a nut shell, why is water such a good solvent for ionic substances?

What is the difference between an electrolyte and a non-electrolyte?

What is meant by a double arrow in a disassociation reaction? Give an example besides the one in the book.

What is the distinction between one arrow and two arrows in a disassociation reaction?

What is meant by a precipitate reaction? Give an example-

What is solubility?

Give the three rules for determining whether a precipitate is formed in a reaction- Last year we learned about Metathesis reactions, we used a different name, what was that name? Show your work for sample exercise 4.3 on page 122-

What is the difference between a molecular equation, a complete ionic equation, and a net ionic equation? Work through an example showing each of the three for the same reaction.

What happens if all the ions in a solution are spectator ions?

Summarize the steps for writing net ionic equations-

Show your work for sample exercise 4.4 below-

Work on problems- 19,20,21,22,23,24,27 which will be due on Thursday