**Chapter 15 Blue Sheet**

What three things can we learn from the example in the book?

Explain in your own words the idea of equilibrium-

What does it mean when we say the equilibrium is dynamic?

What is a reversible reaction?

The reverse reaction is when the \_\_\_\_\_\_\_\_\_\_\_\_ are converted to\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

The forward reaction is when the \_\_\_\_\_\_\_\_\_\_\_\_ are converted to \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

Draw a graph of concentration verses time that demonstrates the idea of equilibrium.

From what side of the reaction is equilibrium reached?

What is the law of mass action? Give a simple mathematical expression of this law-

What do we call the law of mass action?

What is Kc? What does the c stand for?

In general what is the numerator in Kc?

In general what is the denominator in Kc?

Work through sample exercise 15.1-

Using table 15.1 verify Kc for the reaction is a constant value.

In general what is the unit for Kc?

How is Kp different than Kc?

How do Kc and Kp relate?

What is ∆N?

When can Kc and Kp be equal?

What does a **large** value of Kc / Kp mean?

What does a **small** value of Kc / Kp mean?

How does the Kc for the forward reaction relate to the reverse reaction?

How do we find Kc for a reaction that involves more than one step in the reaction?