

3.23 A)

TW:

Same as #2

Sep 16-11:32 AM

3.25 (c) →

$$\begin{aligned} \text{C} &= 7 \times 12 = 84 \\ \text{H} &= 14 \times 1 = 14 \\ \text{O} &= 2 \times 16 = 32 \end{aligned}$$

$$\frac{84}{130} \times 100 = 64.6\%$$

1 → Find total mass  
2 →  $\frac{\text{Wanted Atomic Mass}}{\text{total mass}} \times 100$

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3.33 → (c)

1 → Start with given  
2 → Use AVN to solve for molecules

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345 A EF from % comp

$$\frac{10.4\text{g C}}{12} = \frac{.867\text{ mole}}{.867} = 1$$

$$\frac{27.8\text{g S}}{32} = \frac{.869\text{ mole}}{.867} = 1$$

$$\frac{61.7\text{g Cl}}{35.5} = \frac{1.74\text{ mole}}{.867} = 2$$

$\text{CSCl}_2$

1 → Convert gram out of 100  
2 → Convert to moles  
3 → Find ratio by dividing by lowest #.

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