

TYPE OF  
REACTION

- \_\_\_\_\_ 1.  $\text{Al}_2\text{O}_3 \rightarrow \text{Al} + \text{O}_2$
- \_\_\_\_\_ 2.  $\text{Mg} + \text{HNO}_3 \rightarrow \text{Mg(NO}_3)_2 + \text{H}_2$
- \_\_\_\_\_ 3.  $\text{K} + \text{HOH} \rightarrow \text{KOH} + \text{H}_2$
- \_\_\_\_\_ 4.  $\text{C}_6\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- \_\_\_\_\_ 5.  $\text{Ag} + \text{S} \rightarrow \text{Ag}_2\text{S}$
- \_\_\_\_\_ 6.  $\text{Ca(OH)}_2 + \text{H}_3\text{PO}_4 \rightarrow \text{Ca}_3(\text{PO}_4)_2 + \text{HOH}$
- \_\_\_\_\_ 7.  $\text{Li} + \text{Br}_2 \rightarrow \text{LiBr}$
- \_\_\_\_\_ 8.  $\text{Al} + \text{FeO} \rightarrow \text{Al}_2\text{O}_3 + \text{Fe}$
- \_\_\_\_\_ 9.  $\text{CaO} + \text{P}_2\text{O}_5 \rightarrow \text{Ca}_3(\text{PO}_4)_2$
- \_\_\_\_\_ 10.  $\text{Na}_2\text{CO}_3 + \text{HNO}_3 \rightarrow \text{NaNO}_3 + \text{H}_2\text{CO}_3$

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- \_\_\_\_\_ 11. aluminum + chlorine  $\rightarrow$  aluminum chloride
- \_\_\_\_\_ 12. sodium hydroxide + sulfuric acid  $\rightarrow$  sodium sulfate + water
- \_\_\_\_\_ 13. nickel II chlorate  $\rightarrow$  nickel II chloride + oxygen
- \_\_\_\_\_ 14. butane ( $\text{C}_4\text{H}_{10}$ ) + oxygen  $\rightarrow$  carbon dioxide + water
- \_\_\_\_\_ 15. barium chloride + sodium sulfate  $\rightarrow$  barium sulfate + sodium chloride
- \_\_\_\_\_ 16. zinc + hydrochloric acid  $\rightarrow$  zinc chloride + hydrogen

- \_\_\_\_\_ 17. copper I hydroxide  $\rightarrow$  copper I oxide + water
- \_\_\_\_\_ 18. zinc chloride + ammonium sulfide  $\rightarrow$  zinc sulfide + ammonium chloride
- \_\_\_\_\_ 19. silver nitrate + nickel  $\rightarrow$  nickel II nitrate + silver
- \_\_\_\_\_ 20. sodium chloride + sulfuric acid  $\rightarrow$  sodium sulfate + hydrogen chloride

**Directions : Identify the type of reaction. Complete each of the following equations. Use the information in your packet. Balance each equation.**

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REACTION**

- \_\_\_\_\_ 21. \_\_\_\_ Cd + \_\_\_\_ S  $\rightarrow$
- \_\_\_\_\_ 22. \_\_\_\_ K<sub>2</sub>CO<sub>3</sub>  $\rightarrow$
- \_\_\_\_\_ 23. \_\_\_\_ Mn(ClO<sub>4</sub>)<sub>4</sub>  $\rightarrow$
- \_\_\_\_\_ 24. \_\_\_\_ Al + \_\_\_\_ Cu(NO<sub>3</sub>)<sub>2</sub>  $\rightarrow$
- \_\_\_\_\_ 25. \_\_\_\_ Na + \_\_\_\_ HOH  $\rightarrow$
- \_\_\_\_\_ 26. \_\_\_\_ C<sub>3</sub>H<sub>8</sub> + \_\_\_\_ O<sub>2</sub>  $\rightarrow$
- \_\_\_\_\_ 27. \_\_\_\_ Cu(OH)<sub>2</sub> + \_\_\_\_ HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>  $\rightarrow$
- \_\_\_\_\_ 28. \_\_\_\_ Al(OH)<sub>3</sub>  $\rightarrow$
- \_\_\_\_\_ 29. \_\_\_\_ Pb(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>4</sub> + \_\_\_\_ Ag<sub>2</sub>S  $\rightarrow$
- \_\_\_\_\_ 30. \_\_\_\_ Bi<sub>2</sub>O<sub>3</sub>  $\rightarrow$
- \_\_\_\_\_ 31. \_\_\_\_ FeSO<sub>4</sub> + \_\_\_\_ NH<sub>4</sub>OH  $\rightarrow$

